Name Date

Worksheet 16.1 How much? The amount of chemical change

You can make magnesium oxide by burning magnesium in air. In a laboratory, this is usually done in a crucible with a lid and heating it with a Bunsen burner.

Analysis of results

**1** Record your observations from the experiment, including those which could be evaluated as sources of errors.

**2** Record raw quantitative data in a table. You need to include units and absolute uncertainties   
where appropriate.

**3** Calculate the moles of the magnesium.

**4** Calculate the moles of magnesium oxide formed.

**5** Calculate the theoretical yield of magnesium oxide.

**6** Calculate the % yield.

Evaluation of experiment

**7** Why is the actual yield different to the theoretical yield calculated?

**8** How could you improve your experiment to increase the percentage yield?