Name _____ Date _____

Worksheet 16.1 How much? The amount of chemical change

You can make magnesium oxide by burning magnesium in air. In a laboratory, this is usually done in a crucible with a lid and heating it with a Bunsen burner.

Analysis of results

1 Record your observations from the experiment, including those which could be evaluated as sources of errors.

2 Record raw quantitative data in a table. You need to include units and absolute uncertainties where appropriate.

3 Calculate the moles of the magnesium.

4 Calculate the moles of magnesium oxide formed.

5 Calculate the theoretical yield of magnesium oxide.

6 Calculate the % yield.

Evaluation of experiment

7 Why is the actual yield different to the theoretical yield calculated?

8 How could you improve your experiment to increase the percentage yield?